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Order of Reaction

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Kinetics: Initial Rates  
and Integrated Rate  
Laws **The Rate Law**

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**Kinetics: Chemistry's Demolition Derby - Crash Course**

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*Factors that affect Reaction Rate*

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### AP Chemistry

#### Chapter 14. Chemical Kinetics - 3 -

Instantaneous Rate •

We can plot [C. 4. H.

9. Cl] versus time. •

The rate at any instant in time is called the .

instantaneous rate. •

It is the slope of the straight line tangent to the curve at that instant. •

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Instantaneous rate is different from average rate.

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14.4. Use the equations in the AP Chemistry test booklet to work kinetics problems. Explain the concept of reaction half-life and describe the relationship

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between half-life and rate constant for a first-order reaction.

Use graphical analysis to determine whether the rate law for a reaction is first or second order.

14.5-14.7

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Reaction Rates • All

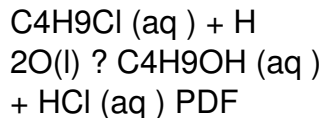
reactions slow down

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over time. •

Therefore, the best indicator of the rate of a reaction is the instantaneous rate near the beginning.



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14.1: Factors that Affect Reaction Rates. chemical kinetics – area of chemistry dealing with speeds/rates of reactions. rates of reactions affected by four factors. concentrations of reactants.

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temperature at which  
reaction occurs.  
presence of a  
catalyst. surface area  
of solid or liquid  
reactants and/or  
catalysts.

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Question 1: A catalyst  
lowers the activation

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energy of a reaction from  $20 \text{ kJ mol}^{-1}$  to  $10 \text{ kJ mol}^{-1}$ . The temperature at which the catalyzed reaction will have the same rate as that of the catalyzed at  $27^\circ \text{C}$  is  $?$

$123^\circ \text{C}$ .  $327^\circ \text{C}$ .  
 $32.7^\circ \text{C} + 23^\circ \text{C}$ .

Correct Option is :

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CHOICE. Choose the  
one alternative that  
best completes the  
statement or answers  
the question. ... 14)

The rate law for a  
reaction is  $\text{rate} = k$   
 $[A][B]^2$  ... The graph  
shown below depicts

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the relationship  
between  
concentration and  
time for the following  
chemical reaction.  
The slope of this ...

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### 14. Chemical Kinetics

- 7 - • Consider the reaction:  $\text{NH}_4^+(\text{aq}) + \text{NO}_2^-(\text{aq}) \rightarrow \text{N}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$  • We measure initial reaction rates. • The initial rate is the instantaneous rate at time  $t = 0$ . • We find this at various initial concentrations of each reactant. • As  $[\text{NH}_4^+]$  doubles with

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concentration and

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time for the following  
chemical reaction.

The slope of this ...

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integrated

zero/first/second

order rate laws,

zero/first/second

order reactions

graphically, & half-life

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Chemical

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It has been found that for a chemical reaction with rise in temperature by  $10^{\circ}\text{C}$ , the rate constant gets nearly doubled.

15. The temperature coefficient of a reaction is the ratio of the rate constants of



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the reaction at two temperatures differing from one another by  $10^{\circ}\text{C}$ .

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